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2: NCERT Solutions - Organic Chemistry, Class 11, Chemistry | EduRev Notes

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It is used for the purification of compounds that are associated with non-volatile impurities or those liquids, which do not decompose on boiling. In other words, distillation is used to separate volatile liquids from non-volatile impurities or a mixture of those liquids that have sufficient difference in boiling points. This method is used to purify a liquid that tends to decompose on boiling. Under the conditions of reduced pressure, the liquid will boil at a low temperature than its boiling point and will, therefore, not decompose. It is used to purify an organic compound, which is steam volatile and immiscible in water. On passing steam, the compound gets heated up and the steam gets condensed to water. After some time, the mixture of water and liquid starts to boil and passes through the condenser. This condensed mixture of water and liquid is then separated by using a separating funnel. Mixture of petrol and kerosene is separated by this method. Glycerol is purified by this method. It boils with decomposition at a temperature of K. At a reduced pressure, it boils at K without decomposition. A mixture of water and aniline is separated by steam distillation. These elements are present in the covalent form in an organic compound. These are converted into the ionic form by fusing the compound with sodium metal. The cyanide, sulphide, and halide of sodium formed are extracted from the fused mass by boiling it in distilled water. In the process, sodium cyanide first reacts with iron II sulphate and forms sodium hexacyanoferrate II. The precipitation of lead sulphide, which is black in colour, indicates the presence of sulphur in the compound. Chemistry of the test The sodium fusion extract is treated with sodium nitroprusside. Appearance of violet colour also indicates the presence of sulphur in the compound. Prussian colour is not formed due to the absence of free cyanide ions. The chemical equation involved in the process can be represented as The traces of nitrogen oxides can also be produced in the reaction, which can be reduced to dinitrogen by passing the gaseous mixture over a heated copper gauge. The dinitrogen produced is collected over an aqueous solution of potassium hydroxide. The volume of nitrogen produced is then measured at room temperature and atmospheric pressure. The nitrogen present in the compound is quantitatively converted into ammonium sulphate. It is then distilled with excess of sodium hydroxide. The ammonia evolved during this process is passed into a known volume of H₂SO₄. The chemical equations involved in the process are The acid that is left unused is estimated by volumetric analysis titrating it against a standard alkali and the amount of ammonia produced can be determined. Thus, the percentage of nitrogen in the compound can be estimated. This method cannot be applied to the compounds, in which nitrogen is present in a ring structure, and also not applicable to compounds containing nitro and azo groups. Discuss the principle of estimation of halogens, sulphur and phosphorus present in an organic compound. In this method, a known quantity of organic compound is heated with fuming nitric acid in the presence of silver nitrate, contained in a hard glass tube called the Carius tube, taken in a furnace. This AgX is then filtered, washed, dried, and weighed. Sulphur, present in the compound, is oxidized to form sulphuric acid. On addition of excess of barium chloride to it, the precipitation of barium sulphate takes place. This precipitate is then filtered, washed, dried, and weighed. Phosphorus, present in the compound, is oxidized to form phosphoric acid. By adding ammonia and ammonium molybdate to the solution, phosphorus can be precipitated as ammonium phosphomolybdate. Explain the principle of paper chromatography. This paper contains water trapped in it, which acts as the stationary phase. On the base of this chromatography paper, the solution of the mixture is spotted. The paper strip is then suspended in a suitable solvent, which acts as the mobile phase. This solvent rises up the chromatography paper by capillary action and in the procedure, it flows over the spot. The components are selectively retained on the paper according to their differing partition in these two phases. The spots of different components travel with the mobile phase to different heights. The paper so obtained shown in the given figure is known as a chromatogram. Why is nitric acid added to sodium extract before adding silver nitrate for testing halogens? That is, if any nitrogen and sulphur are present in the form of NaCN and Na₂S, then they are removed. The chemical equations involved in the reaction are represented as Question Explain

the reason for the fusion of an organic compound with metallic sodium for testing nitrogen, sulphur and halogens. For their detection, they have to be first converted to ionic form. This is done by fusing the organic compound with sodium metal. The chemical equations involved in the test are Carbon, nitrogen, sulphur, and halogen come from organic compounds. Name a suitable technique of separation of the components from a mixture of calcium sulphate and camphor. In this process, the sublimable compound changes from solid to vapour state without passing through the liquid state. Camphor is a sublimable compound and calcium sulphate is a non-sublimable solid. Hence, on heating, camphor will sublime while calcium sulphate will be left behind. Explain, why an organic liquid vaporises at a temperature below its boiling point in its steam distillation? Give reason for your answer. This is because the chlorine atoms are covalently bonded to carbon in CCl_4 . Why is a solution of potassium hydroxide used to absorb carbon dioxide evolved during the estimation of carbon present in an organic compound? Hence, carbon dioxide reacts with potassium hydroxide to form potassium carbonate and water as Thus, the mass of the U-tube containing KOH increases. From its mass, the percentage of carbon in the organic compound can be estimated. Why is it necessary to use acetic acid and not sulphuric acid for acidification of sodium extract for testing sulphur by lead acetate test? Hence, it is necessary to use acetic acid for acidification of sodium extract for testing sulphur by lead acetate test. Calculate the masses of carbon dioxide and water produced when 0. Percentage of hydrogen in organic compound is 4. Thus, 2 g of hydrogen is contained in 18 g of water. A sample of 0. The ammonia evolved was absorbed in 50 mL of 0. The residual acid required 60 mL of 0. Find the percentage composition of nitrogen in the compound. Calculate the percentage of chlorine present in the compound. Thus, mass of chlorine in 0. In the estimation of sulphur by Carius method, 0. Find out the percentage of sulphur in the given compound.

3: Chemistry - General Knowledge Questions and Answers

Chapter 2 - Structure of Atom Structure of atom class 11 questions and answers are intended for students of class 11 who are looking to improve and test their knowledge in the subject. When students shift to the 11th standard from class 10, they usually tend to experience a huge gap in the syllabus of class 11th.

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wastes; green chemistry as an alternative tool for reducing pollution, strategy for control of environmental pollution. Total Periods 60 Micro-chemical methods are available for several of the practical experiments. Wherever possible such techniques should be used. Basic Laboratory Techniques Periods 2

1. Cutting glass tube and glass rod
2. Bending a glass tube
3. Drawing out a glass jet
4. Boring a cork

B. Determination of melting point of an organic compound. Determination of boiling point of an organic 3. Crystallization involving impure sample of any one of the following: Alum, copper sulphate, Benzoic acid. Determination of pH of some solutions obtained from fruit juices, solutions of known and varied concentrations of acids, bases and salts using pH paper or universal indicator. Comparing the pH of solutions of strong and weak acid of same concentration. Study the pH change in the titration of a strong acid with a strong base using universal indicator. Chemical Equilibrium Periods 4 One of the following experiments: Quantitative Estimation Periods 16 Using a chemical balance. Preparation of standard solution of oxalic acid. Determination of strength of a given solution of sodium hydroxide by titrating it against standard solution of oxalic acid. Preparation of standard solution of sodium carbonate. Determination of strength of a given solution of hydrochloric acid by titrating it against standard sodium carbonate solution. Insoluble salts excluded b Detection of nitrogen, sulphur, chlorine, in organic compounds. Project Periods 10 Scientific investigations involving laboratory testing and collecting information from other sources. A few suggested projects Checking the bacterial contamination in drinking water by testing sulphide ions. Study of the methods of purification of water. Testing the hardness, presence of iron, fluoride, chloride etc. Study of the acidity of different samples of the tea leaves. Determination of the rate of evaporation of different liquids. Study of the effect of acids and bases on the tensile strength of fibers. Analysis of fruit and vegetable juices for their acidity. Any other investigatory project, which involves about 10 periods of work, can be chosen with the approval of the teacher. CBSE has regional offices which have been set up to handle regional matters and also solve the queries of schools and students coming in their region. In this regard CBSE has issued the following important circular for the benefit of students and parents. You are aware that Central Board of Secondary CBSE has requested the affiliated schools to organize events in schools and educate the students on the impact of corruption on our society and how to eradicate it. CBSE has issued the following circular which you should read as there During this time many chapters will be modified, deleted and new chapters might also be added based on the latest education trend Easier Question Papers for class 10 and class 12 this year CBSE has recently released the sample question papers for class 10 and class 12 board exams for The sample papers were released last week with marking schemes which the students can download in pdf format and practise them to understand the type of questions which are going to come in the board exams. CBSE has also released the marking schemes for all the papers to help students to Board Exams Distribution of subjects for February March and April CBSE has issued a circular providing details regarding the board exams to be held for class 10 and class 12 in The circular gives details of the subjects for which the exams will be held in February , March and April The sample papers have been issued along with the marking schemes. These sample papers are made exactly in the same pattern as the papers which come in the examinations. Usually all sample papers are of marks and are issued to students atleast couple of months prior to the examinations so

4: CBSE Class 11 Chemistry- Chapter Wise Important Questions Free Download

The NCERT Solutions for Class 11 Chemistry includes answers to questions ranging across various chapters such as the Structure of Atom, States of Matter, Thermodynamics, Equilibrium, Redox Reactions, Hydrogen, and others.

Barium plus a Chlorine producer Blue Copper plus a Chlorine producer Purple Silver Aluminium or Magnesium powder Oranges, reds and whites are usually made through incandescence. Blues and greens are usually made through luminescence. What Is an Alloy? Alloys are mixtures containing at least one metal. Take iron - whilst extremely strong, it is also very brittle Add in a little carbon and you make steel - an alloy with the strength of iron but it not brittle. Alloys contain atoms of different sizes, which makes it more difficult for the atoms to slide over one another. This makes alloys harder than the pure metal. Certain mixtures are even more impressive. Mix nickel and titanium and you get Nitinol, a smart alloy used to make spectacle frames. Nickel-Iron alloys are common in meteorites. How Does a Match Light? Match heads are made using phosphorus - a highly combustible element - that catches fire due to the friction caused when striking the match. Safety matches are slightly different. They will only light if you strike them using the surface on the side of the box. In this case, the match head contains potassium chlorate - an accelerant that speeds up the reaction. The rough side of the box contains most of the phosphorus. Bring the two together and add the heat generated by friction, and you have a flame. Waterproof matches have a thin coating of wax over the whole match. This is removed when striking the head against the box, exposing the phosphorus. This allows the match to catch. To give you enough time to move the match to whatever you wish to light, most matchsticks are treated with paraffin candle wax. The first friction match was invented in by English chemist, John Walker. The earliest match is thought to have arisen in China in AD. These were nothing more than sticks impregnated with sulphur. Bubbles in fizzy drinks can only form at points called nucleation sites - these are sharp edges or bits of dirt or grime that helps the release of carbon dioxide gas. A mento is actually not as smooth as it appears. Under a microscope you can see that there are millions of tiny craters on the surface. Each of these provides a nucleation site for carbon dioxide gas to form. Diet coke works best because the surface tension in the drink is much lower than regular coke - this allows bubbles to form more easily. This is due to the substitution of sugar with the sweetener aspartame. What is the Ozone Layer? The ozone layer is a massive shield that surrounds the Earth, 50kms above the surface of the planet. Ozone is a special molecule of oxygen: It is up to 20kilometres thick and most of this gas is found in the stratosphere. Ozone gases are our protection against UVB radiation. This damaging radiation is emitted by the Sun and is extremely dangerous. The ozone hole is largely over the Antarctic and is between 21 and 24 million square kilmoetres in size. The hold is caused by ozone reacting with CFCs - pollutants used in refrigeration. The largest recorded ozone hole occurred in at What is your favourite chemistry experiment mentioned in this hub?

5: 11th Class Chemistry Chapter 3 Gases Short Question Answers Fsc Part 1 Easy Notes

Solve important practice questions of Class 11 Chemistry. Prepare with important problems for exams. Access questions bank with more than 80, questions & solutions.

Name the first element of periodic element. What is its electronic configuration? Hydrogen is the first element in the periodic table and also the lightest element known. Name the three isotopes of hydrogen. Which isotope of hydrogen is radioactive? How is dihydrogen prepared commercially? By electrolysis of acidified water. What is the importance of heavy water with regard to nuclear power generation? It is used as a moderator and helps to control the nuclear reaction. Name the compound whose electrolysis in aqueous state, give high purity. Name the element used in fuel cell for generating electricity. What is a zeolite? Name the hydrides which have high potential for hydrogen storage. Give an example of electron deficient covalent hydride. Write two uses of interstitial hydrides. What type of elements form interstitial hydrides? Name the groups in d-block elements which do not form metallic hydrides. It is used as an antiseptic. Which isotope of hydrogen does not have neutron? What is the use of hydrogen in the manufacture of Vanaspati Ghee? It is also called hydrogenation. Name the phenomenon of adsorption of hydrogen on metal surface. List the uses of dihydrogen?

6: Top Ten Chemistry Questions and Answers | Owlcation

All Chapter 1 - Some Basic Concepts of Chemistry Exercises Questions with Solutions to help you to revise complete Syllabus and boost your score more in examinations. Free PDF download of NCERT Solutions for Class 11 Chemistry Chapter 1 - Some Basic Concepts of Chemistry solved by Expert Teachers as per NCERT (CBSE) textbook guidelines.

Lighter gases diffuse more rapidly than heavier gases. Actually in kinetic molecular theory of gas we suppose that there are forces of attraction. At high temperature attractive forces become less dominant and gases behave ideally. The actual volume of O₂ gas at room temperature and 1 atm. Pressure is negligible as compared to volume occupied by one mole of this gas. But this actual volume is not negligible at high pressures. The actual volume of the gas molecules is negligible as compared to the volume of vessel. When the pressure is increased for O₂, then molecules come close to each other. Force of attraction start dominating actual volume remains no more negligible. Hydrogen and helium are ideal at room temperature but SO₂ and Cl₂ are non-ideal. How do you explain it? H₂ and He have very low boiling points. So at room temperature they are far away from their boiling points. At room temperature the attractive forces are absent. So they behave ideally. At room temperature they are not far away from their boiling points. Sufficient attractive forces are present at room temperature. So they are non-ideal. How the critical temperature is an essential criteria to be considered for the liquefaction of gases? Gases can be liquefied by increasing the pressure and decreasing the temperature. The temperature of a gas should reach the critical temperature on or below that to make the gas liquid. Gas can never be liquefied how much the pressure is applied if the gas is above the critical temperature. Why the non-polar gases like H₂ and He have a very low critical temperatures while polar gases like NH₃ and SO₂ have critical temperatures sufficiently above room temperature? H₂ and He are consisted of small-sized molecules and have low polarities. They have least forces of attractions among themselves. NH₃ and SO₂ have attractive forces even at room temperature. In order to liquefy them their temperature may be above the room temperature. In Joule-Thomson effect sudden expansion of the gas molecules needs energy. In the compressed state there are sufficient attractive forces among the molecules of the gas. During sudden expansion the energy is required to overcome the inter molecular attractions. Moreover the molecules need extra energy to run away in vacuum. These compressed gases are suddenly allowed to expand through a nozzle. Why the gases deviate from ideal behavior at high pressure and low temperature? When the pressure of the gases are high, collisions become more frequent and force of attraction are created. Moreover the actual volume of the gas molecules are no more negligible as compared to the volume of the vessels. So two postulates of this theory are faulty. Water vapour does not behave ideally at K. Due to this reason water vapour behaves non-ideally at K. Give two causes for deviation of gases from ideality? The causes are due to two faulty assumptions. These two postulates are correct when the temperature is high or pressure is low. SO₂ is comparatively non-ideal at K but behaves ideally at K. SO₂ gas is close to its boiling point at K. So, at K attractive forces are dominating and make the gas non-ideal. But when the temperature of the gas is K, then forces of attractions are less dominant and gas behaves ideally. These parameters are very important for the real gases. Why the pressure correction is done by van der Waals? The pressure is exerted due to collisions on the walls of vessel. Attractive forces decrease the intensity of collisions. So the observed pressure on the walls is less than the ideal pressure. Lessened pressure is added in observed pressure to get ideal pressure. Why the volume correction is done by van der Waals? The molecules of a gas do occupy certain volume which is not available to the gas molecules in the vessel.

7: NCERT Solutions for Class 11 Chemistry

NCERT Solutions Class 11 Chemistry Chemistry Lab Manual Chemistry Sample Papers. NCERT TEXTBOOK QUESTIONS SOLVED. Question 1. A liquid is in equilibrium with its.

8: General Knowledge Questions Answers Class 12 & GK Quiz for Students - Entrance Exam Information

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