

1: Vapour Pressure - Physical Properties Sources Index (PPSI): Details

Computer Aided Data Book of Vapor Pressure Second Edition, by Shuzo Ohe, Ph.D. This second edition is a major revision of the original, which is cited in many well known professional books.

Phase Equilibrium Data Last Updated on Tue, 07 May Extractive Distillation For a binary mixture , pressure and temperature fix the equilibrium vapor and liquid compositions. Thus, experimental data are frequently presented in the form of tables of vapor mole fraction y and liquid mole fraction x for one constituent over a range of temperature T for a fixed pressure P or over a range of pressure for a fixed temperature. A compilation of such data, mainly at a pressure of For application to distillation a nearly isobaric process , as shown in Figs. In most binary systems, one of the components is more volatile than the other over the entire composition range. This is the case in Figs. For some binary systems, one of the components is more volatile over only a part of the composition range. Two systems of this type, ethyl acetate-ethanol and chloroform-acetone, are shown in Figs. Figure shows that for two binary systems chloroform is less volatile than acetone below a concentration of 66 mole percent chloroform and that ethyl acetate is less volatile than ethanol below a concentration of 53 mole percent ethyl acetate. Above these concentrations, volatility is reversed. Such mixtures are known as azeotropic mixtures, and the composition in which the reversal occurs, which is the composition in which vapor and liquid compositions are equal, is the azeotropic composition, or azeotrope. The azeotropic liquid may be homogeneous or heterogeneous two immiscible liquid phases. Many of the binary mixtures of Table form homogeneous azeotropes. Non-azeotrope-forming mixtures such as benzene and toluene in Figs. By contrast, simple distillation of azeotropic mixtures will at best yield the azeotrope and one essentially pure species. The distillate and bottoms products obtained depend upon the feed composition and whether a minimum-boiling azeotrope is formed as with the ethyl acetate-ethanol mixture in Fig. For example, if a mixture of 30 mole percent chloroform and 70 mole percent acetone is fed to a simple distillation column, such as that shown in Fig. An example of heterogeneous-azeotrope formation is shown in Fig. At liquid compositions between 0 and 3 mole percent butanol and between 40 and mole percent butanol, the liquid phase is homogeneous. Phase splitting into two separate liquid phases one with 3 mole percent butanol and the other with 40 mole percent butanol occurs for any overall liquid composition between 3 and 40 mole percent butanol. For mixtures containing more than two species, an additional degree of freedom is available for each additional component. Thus, for a four-component system, the equilibrium vapor and liquid compositions are only fixed if the pressure, temperature, and mole fractions of two components are set. Representation of multicomponent vapor-liquid equilibrium data in tabular or graphical form of the type shown earlier for binary systems is either difficult or impossible. Instead, such data, as well as binary-system data, are commonly represented in terms of K values vapor-liquid equilibrium ratios , which are defined by and are correlated empirically or theoretically in terms of temperature, pressure, and phase compositions in the form of tables, graphs, and equations. Rarely is distillation used on a large scale if the relative volatility is less than 1.

2: Beryllium fluoride - Wikipedia

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3: Vapor Pressure Data

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4: Vapor-liquid equilibrium data. | University of Texas Libraries

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7: Antoine Constants - Physical Properties Sources Index (PPSI): Details

More extensive and detailed vapor pressure data on selected important substances appear in other tables in this section of the Handbook. Computer Aided Data Book.

8: 1,2-Dibromoethylene, (1E)- | C₂H₂Br₂ - PubChem

Computer Aided Data Book of Vapor Pressure. Number of compounds not specified. www.enganchecubano.comix A: Sources of vapour pressure data Presented list of sources of vapour pressure data of is definitely not complete.s-ohe.

9: Diethyl adipate | C₁₀H₁₈O₄ - PubChem

(1) Ohe S; Computer Aided Data Book of Vapor Pressure. Tokyo, Japan: Data Book Publ Co () (2) Stephen H, Stephen T, ed; Solubilities of inorganic and organic compounds. Vol 1.

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