

CONDUCTIVITIES AND VISCOSITIES IN PURE AND IN MIXED SOLVENTS.

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Excerpt from Conductivities and Viscosities in Pure and in Mixed Solvents: Radiometric Measurements of the Ionization Constants of Indicators, Etc An examination of the literature bearing on viscosity shows that, in general, only the salts of the three metals mentioned above are known to lower the viscosity of water.

Kinematic Viscosity Measurements of Mixed solvents of 0. The density values decreased with decrease in the mole fraction of THF for the binary system tables 3,4,9, Density is a measure of solvent- solvent and ion-solvent interactions Thirumaran et al, The decrease indicates formation of less dense products. The densities and viscosities of the systems with Mg ClO₄ 2 increased with concentration which suggests a moderate strong associative nature in which the solute Mg ClO₄ 2 tends to attract the solvent molecules, which indicates an increase in solvent- solvent and solute- solvent interactions which leads to volume shrinkage. Thirumaran and Satish, have also made such observations. The observed decrease in densities and viscosities with increase in temperature is attributed to greater thermal excitation and reduction of attractive forces between the ions. Rodriguez et al, , Nageshewar et al, and Motin, , have reported such. Conductivity The conductivity of an electrolyte solution depends on the concentration of the ionic species. The higher the dielectric constant and the lower the viscosity of a solvent, the higher the conductivity of the solution. Such observations have been reported by Matsuda and Satake , and Saito et al, Such observation has also been made by Corkum et al, The results show that ion-solvent and solventsolvent interactions contribute to the improvement of conductance. Matsuda et al, have reported such. A net increase in ion conductivity is achieved when solvents of varying properties are mixed. Increase in THF content or low AN content result in a very low dielectric medium where the effect of ion pair formation outweighs that of low viscosity; thus, ion conductivity drops with the increase in THF content. Therefore, the maximum in ion conductivity versus solvent composition is actually the result of the compromise between the effects of the dielectric constant and viscosity. Such a compromise illustrates the superiority of mixed solvent over single solvent electrolytes. The trend in conductance can be described by the viscosity behavior of the electrolytic media. Jones- Doles A and B Coefficients: The B coefficients for the system were positive for the entire fractions ,indicating strong ion interactions with the solvent molecules. The difference in concentrations is the only factor that produces cell potential in the concentration cells Anusiem, This shows the dependence of cell potential on concentration. The results of the physico-chemical parameters show that the values are intermediate between those of the pure solvents. This work shows that the mixing of organic solvents improves the physico properties of the pure solvents, thereby making them suitable for use as electrolytes in Magnesium ion batteries. Frankfurt , 12, 1, Rodriguez, A, Canosa J. Saito, T, Ikeda, H. Journal of Applied Electrochemistry, 6, 85

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